

KEY

Name: _____
Period: _____ Date: _____

Chemistry: Ionic Binary Compounds: Single-Charge Cations

Write the name of each of the following compounds.

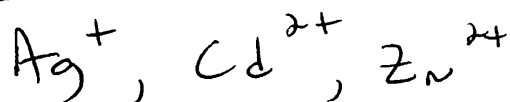
1. Na_2S
2. Al_2O_3
3. NaCl
4. RbI
5. ZnBr_2
6. AgCl
7. BN weird compound.
8. BaF_2
9. Sr_3N_2
10. MgCl_2

1. Sodium sulfide
2. Aluminum oxide
3. Sodium chloride
4. Rubidium iodide
5. Zinc bromide
6. Silver chloride
7. Boron nitride
8. Barium fluoride
9. Strontium nitride
10. Magnesium chloride

Write the chemical formula for each of the given names.

11. magnesium nitride
12. calcium oxide
13. silver fluoride
14. beryllium chloride
15. potassium iodide
16. aluminum chloride
17. zinc oxide
18. barium bromide
19. lithium nitride
20. potassium sulfide

11. Mg_3N_2
12. CaO
13. AgF
14. BeCl_2
15. KI
16. AlCl_3
17. ZnO
18. BaBr_2
19. Li_3N
20. K_2S



← Always

Chemistry: Ionic Binary Compounds: Multiple-Charge Cations

Write the name of each of the following compounds.

1. CuF 1. copper (I) fluoride
2. CuF_2 2. copper (II) fluoride
3. Cr_2O_3 3. chromium (III) oxide
4. PbI_2 4. lead (II) iodide
5. PbCl_4 5. lead (IV) chloride
6. CrO_3 6. chromium (VI) oxide
7. AuBr 7. gold (I) bromide
8. NiO 8. Nickel (II) oxide
9. V_2I_6 9. vanadium (III) iodide
10. SnO_2 10. tin (IV) oxide

Write the chemical formula for each of the given names.

11. manganese (VII) oxide 11. Mn_2O_7
12. niobium (V) chloride 12. NbCl_5
13. titanium (III) phosphide 13. Ti_3P_2
14. palladium (IV) sulfide 14. PdS_2
15. platinum (II) fluoride 15. PtF_2
16. osmium (III) oxide 16. Os_2O_3
17. iridium (IV) nitride 17. Ir_3N_4
18. cobalt (II) chloride 18. CoCl_2
19. iron (III) sulfide 19. Fe_2S_3
20. gold (III) iodide 20. AuI_3