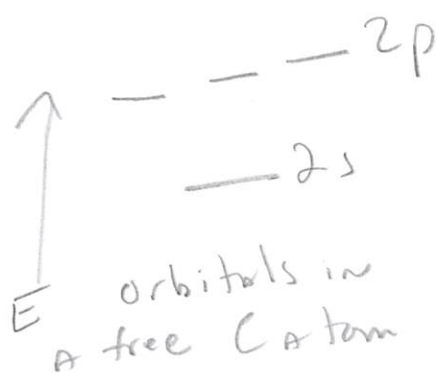
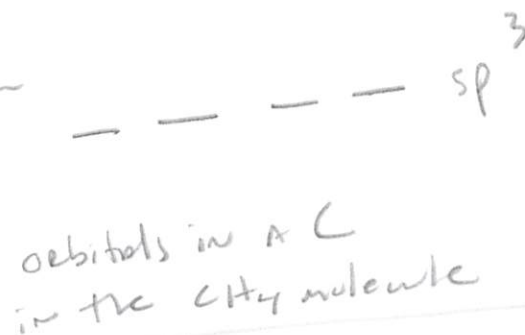


sp^3 hybridization



hybridization
⇒

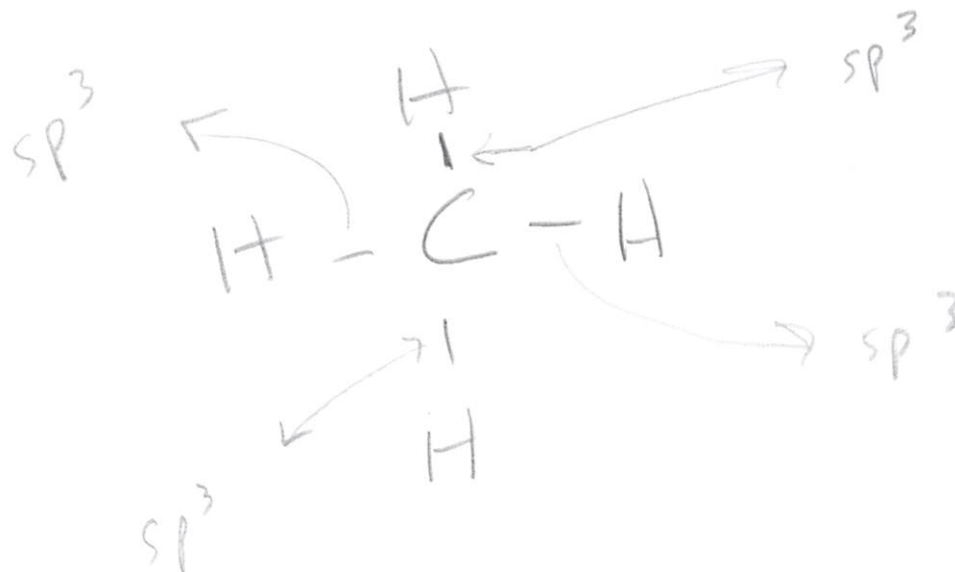


★ one 2s orbital and three 2p orbitals
become four sp^3 hybridized orbitals



AND NOW CAN MAKE 4 BONDS

such as

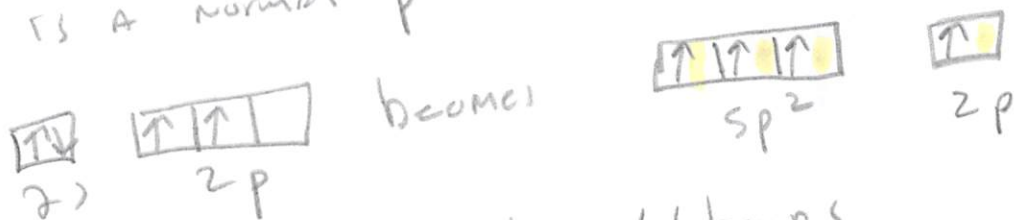


sp² hybridization



★ one 2s orbital and two 2p orbitals become three sp² hybridized orbitals

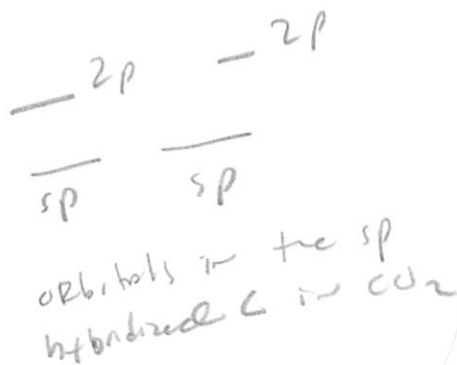
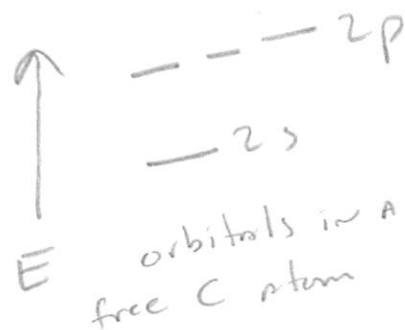
★ notice that one of the p orbitals of the atom HAS NOT been used and is a normal p orbital



and can make 4 bonds
such as



sp hybridization



- ★ one 2s and one 2p orbitals become two sp hybridized orbitals
- ★ notice that two of the p orbitals of the atom HAVE NOT been used and are normal p orbitals



and can make 4 bonds
such as

