

### **Chapter 6 Review**

1. Be able to define: electronegativity, ionization energy, cation, anion, atomic radius, transition metal, representative elements, periodic law, group, period
2. Be able to describe Mendeleev's role in the periodic table's development.
3. Metals/Non-metals dividing line
4. Infer information from the atomic number of an element.
5. Be able to do electron configuration.
6. Noble gas electron configuration—why is it special?
7. Understand trends in atomic radius.
8. What would increase the “shielding” effect?
9. Why does atomic size increase within a group?
10. Cation/Anions—definitions/charge/what groups are likely to form which?
11. Ionization energy definition/difference between 1<sup>st</sup>, 2<sup>nd</sup> 3<sup>rd</sup> Ionization energy
12. Symbols for ions/correct charge based on Group #
13. Electronegativity trends—highest/lowest
14. Size of parent atom compared to its ion
15. Evaluate “Groups” as to whether they are likely to gain or lose electrons and how many
16. What leads to a decrease in ionization energy as you go down a group?

**DON'T UNDERESTIMATE THE DIFFICULTY OF ANSWERING QUESTIONS ABOUT TRENDS!**

**YOU MUST MEMORIZE TRENDS AND UNDERSTAND THE REASONS FOR THE TREND**