To find the $[H^+]$ from pH

example

$pH = 5.00$

$-\log[H^+] = 5.00$

so.. $\log[H^+] = -5.00$

on calculator

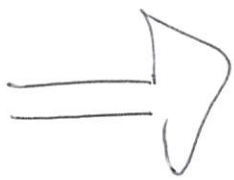
+/− 5.00 10^x

↑

0.00001 =

1.0×10^{-5} M

To find pH from concentration you would NOT use 10^x



To Find pH from Concentration

use $\log_{10}$ button

→ use ( ) parenthesis if the M is in SCINOT

→ if NOT in SCINOT, parenthesis are NOT necessary

$M = 6.35 \times 10^{-10}$ mol/L

(6.35 EE +/- 10) $\log_{10}$ →

-9.2 so

$-(-9.2) = 9.2 = pH$

$\log_{10}$

for FIND the pH

ADD A NEGATIVE SIGN
to the value

AND 10^{x}

for find the concentration.

SUMMARY