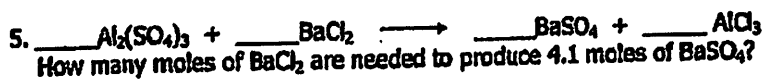
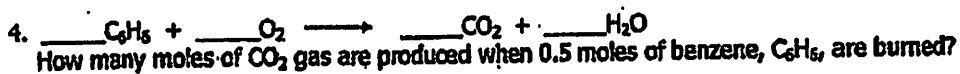
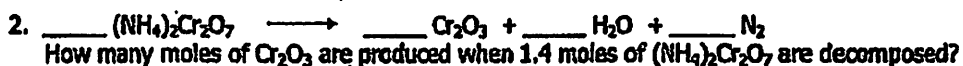
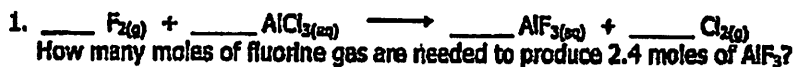


Mole-Mole Stoichiometry Practice

Name:

Hour:

Balance the following equations, then answer the question that follows.



Mole-Mass Stoichiometry Practice

Name:

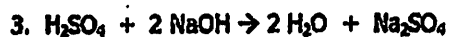
Hour:



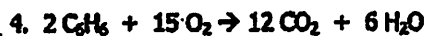
How many moles of fluorine gas are needed to produce 50 g of AlF_3 ?



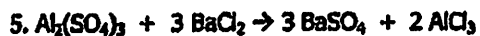
How many grams of Cr_2O_3 are produced when 1.2 moles of $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$ are decomposed?



How many moles of NaOH are needed to neutralize 520 grams of H_2SO_4 ?



How many moles of CO_2 gas are produced when 85 grams of benzene, C_6H_6 , are burned?



How many grams of BaCl_2 are needed to produce 4.1 moles of BaSO_4 ?

Mass-Mass Stoichiometry Practice

Name:

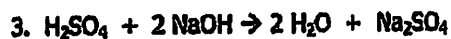
Hour:



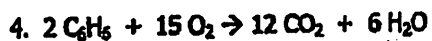
How many grams of fluorine gas are needed to produce 100.00g of AlF_3 ?



How many grams of Cr_2O_3 are produced when 25.00g of $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$ is decomposed?



How many grams of NaOH are needed to neutralize 15.00g of H_2SO_4 ?



How many grams of CO_2 gas can be produced of 156.24g of benzene, C_6H_6 , is burned?



How many grams of BaCl_2 are needed to produce 25.00g of BaSO_4 ?

Mole-Particle Stoichiometry Practice

Name: _____

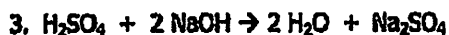
Hour: _____



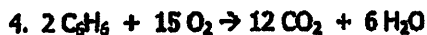
How many moles of fluorine gas are needed to produce 1.5×10^{24} molecules of Cl_2 ?



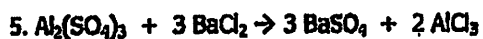
How many molecules of H_2O are produced when 7.8×10^{23} formula units of $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$ are decomposed?



How many moles of NaOH are needed to neutralize 5.20×10^{24} particles of H_2SO_4 ?



How many molecules of CO_2 gas are produced when 8.5 moles of benzene, C_6H_6 , are burned?



How many grams of BaCl_2 are needed to produce 4.1×10^{23} particles of BaSO_4 ?