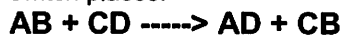


CHEMISTRY REPLACEMENT REACTION WORKSHEET

DISCRIPTION

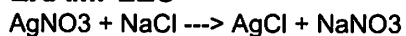
During double replacement, the cations and anions of two different compounds switch places.



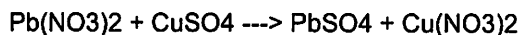
REACTION GUIDELINE

1. It is important that the formulas of the products be written correctly. If they are correct, balancing the equation is a simple task; if not, the equation will probably never balance.
2. In these reactions, there is never a change in oxidation state (in other words, the charges stay the same).

EXAMPLES



Silver nitrate + sodium chloride $\rightarrow$ silver chloride + sodium nitrate



Lead (II) nitrate + copper (II) sulfate $\rightarrow$ lead (II) sulfate + copper (II) nitrate

DOUBLE REPLACEMENT PRACTICE REACTIONS

For each reaction predict the products and balance the equation. State the reaction in chemical formulas and in symbols. For example:



Silver nitrate + sodium chloride $\rightarrow$ silver chloride + sodium nitrate

